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Antimicrobials: Mechanism, classification, Potassium permanganate, Boric acid, Hydrogen peroxide*, Chlorinated lime*, Iodine and its preparations

- An antimicrobial is an agent that kills microorganisms or stops their growth.
- The most common targets for antimicrobial drug actions fall into 5 basic categories:
 - Inhibition of Cell Wall Synthesis
 - Inhibition of Protein Synthesis
 - Inhibition of Nucleic Acid Synthesis
 - Effects on cell membrane sterols (antifungal agents)
 - Inhibition of unique metabolic steps.
- Control of Microbial Growth:**
 - Sterilizing Agents:**
 - Disinfectants:** Disinfectants are antimicrobial agents that are applied to the surface of non-living objects to destroy microorganisms that are living on the objects. Disinfection does not necessarily kill all microorganisms, especially resistant bacterial spores; it is less effective than sterilization.
 - Antiseptics:** Antiseptics (from Greek anti: "against" and séptikos: "putrefactive") are antimicrobial substances that are applied to living tissue/skin to reduce the possibility of infection, sepsis, or putrefaction.
 - Chemotherapeutics:** The chemical agents which is used in the clinical application of antimicrobial agents to treat infectious disease.
- Medicines with an antimicrobial activity are divided into two groups:**
 - Non-selective antimicrobial agents, causes most destructive effect on the majority of microorganisms (antiseptics and disinfectants).
 - Selective antimicrobial drugs (chemotherapeutic agents)
- Requirements for antiseptics and disinfectants**
 - Must have a broad spectrum of action;
 - Rapid onset of action;
 - Should have a small latency period;
 - Should have a high activity;
 - Must be chemically resistant;
 - High availability and low cost;
 - Lack of local irritant or allergic effects on tissues;
 - Minimal absorption from the place of their application;
 - Low toxicity.



Classification of Antiseptics and Disinfectants (according chemical structure)

1. Inorganic substances

- | | | | |
|---|---|---|--|
| (i) Halogens | (ii) Oxidizing agents | (iii) Acids and alkalis | (iv) Metallo salts |
| o Iodine
(2%, 3%, 5% alcohol solution) | o Hydrogen peroxide
o Potassium permanganate | o Boric acid
o Salicylic acid
o Solution of ammonia | o Silver nitrate
o Copper sulfate
o Zinc sulfate
o Zinc oxide |
| o Povidon-Iodine
o Lugol's solution
o Chlorhexidine | | | |

2. Organic substances

- | | | | |
|--|----------------------|--|---------------------------------------|
| (i) Aldehydes | (ii) Alcohols | (iii) Phenol derivatives | (iv) Dyes |
| o Formaldehyde
o Glutaraldehyde
o Hexamethylenetetramine | o Ethyl alcohol | o Phenol
o Cresol
o Resorcinol
o Thymol | o Methylene blue
o Brilliant green |

Textbook of Inorganic Pharmaceutical and Medicinal Chemistry

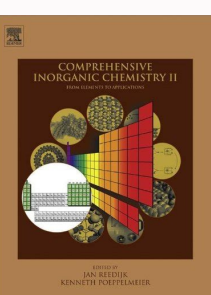
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Review

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- **Acidifiers** are inorganic chemicals that either produce or become acid.
- These chemicals increase the level of gastric acid in the stomach when ingested, thus decreasing the stomach pH.
- There are many types of acidifiers but the main four types are:
 - Gastric acidifiers*, used in controlling pH in stomach.
 - Urinary acidifiers*, used in controlling pH in urine.
 - Systemic acidifiers*, used in controlling pH in all the parts of body.
- **ACHLORHYDRIA:** In patients suffering from achlorhydia, there is deficient secretion of HCl in stomach. In such cases acidifiers are useful in providing the necessary acidity for the proper digestion of food. Systemic acidifiers are those which, when given usually by injection, act by reducing the alkali reserve in the body and are also useful in reducing metabolic alkaloids.

❖ Ammonium chloride

Molecular formula	NH ₄ Cl
Molar mass	53.49 g/mol
Synonym	Sal ammoniac



Solutions of ammonium chloride are mildly acidic.

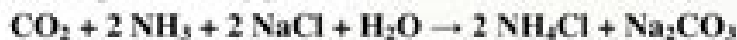
Physical Properties

Appearance: White solid, hygroscopic **Odor:** Odourless **Taste:** Cooling saline **Density:** 1.5274 g/cm³

Melting point: 338° C (decomposes, sublimes) **Solubility:** free soluble in water and glycerol, Sparingly soluble in alcohol

Preparation:

- Ammonium chloride prepared through the *Solway process*:



- Ammonium chloride is prepared commercially by combining ammonia (NH₃) with either hydrogen chloride (gas) or hydrochloric acid (water solution): **NH₃ + HCl → NH₄Cl**

Reactions

- Ammonium chloride appears to sublime upon heating but actually decomposes into ammonia and hydrogen chloride gas: **NH₄Cl → NH₃ + HCl**

- Ammonium chloride reacts with a strong base, like sodium hydroxide, to release ammonia gas:



- Similarly, ammonium chloride also reacts with alkali metal carbonates at elevated temperatures, giving ammonia and alkali metal chloride: **2 NH₄Cl + Na₂CO₃ → 2 NaCl + CO₂ + H₂O + 2 NH₃**

Assay

- Dissolve 1.000 g of Ammonium chloride in 20 ml of water and add a mixture of 5 ml of formaldehyde solution, with few drops of phenolphthalein solution. After 1 min to 2 min, titrate slowly with 1M sodium hydroxide.

- **1 ml of 1M sodium hydroxide is equivalent to 53.49 mg of NH₄Cl.**

Uses:

- Ammonium chloride is used as an expectorant in cough medicine. Its expectorant action is caused by irritative action on the bronchial mucosa, which causes the production of excess respiratory tract fluid, which presumably is easier to cough up.
- Ammonium salts are an irritant to the gastric mucosa and may induce nausea and vomiting.
- Ammonium chloride is used as a systemic acidifying agent in treatment of severe metabolic alkalosis.
- The main application of ammonium chloride is as a nitrogen source in fertilizers.
- Ammonium chloride is used as a flux in preparing metals to be tin coated, galvanized or soldered.

- **Dose:** 1 to 2 gm (As systemic acidifier); 0.3 to 0.5 gm (Expectorant)

- **Storage:** Store in highly closed container.

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